

Original Research

Preparation and Adsorption Properties of Chitosan-Modified Magnetic Nanoparticles for Removal of Mo (VI) Ions

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Abstract

In this study, the new magnetic nanoparticles modified with chitosan were prepared by reverse suspension cross-linking method to solve the problem of excess molybdenum in mining wastewater. The materials were employed for the adsorption of Mo (VI) from water. By adjusting the pH value of the water sample, the adsorption process of molybdenum was completed after oscillating the adsorbed molybdenum for a period of time between composite nanometer adsorbent and Mo (VI) at a certain speed. The concentration of molybdenum was determined by thiocyanate spectrophotometry. The result showed that under the condition of pH = 4.0, the adsorption effect of Mo (VI) on chitosan-modified magnetic nanoparticles was the best. The adsorption process accorded with the Langmuir adsorption isotherm model, and the high fitting degree with the quasi-second-order adsorption kinetic model accorded with chemical adsorption. Theoretical maximum adsorption capacity was 35.54 mg/g. The adsorption of Mo (VI) on chitosan-modified magnetic nanoparticles was still over 90% after 3 regenerations. Removal of Mo (VI) by chitosan-modified magnetic nanoparticles is quick and efficient. Its reusability is good. Its magnetic property is strong and it is convenient for solid-liquid magnetic separation. More importantly, our results provide a sustainable approach for removal of Mo (VI), chitosan-modified magnetic nanoparticles have more potential to be exploited and utilized for metal adsorption.

Keywords: molybdenum, magnetic nanoparticles, chitosan, adsorption, water treatment

Introduction

Molybdenum is a necessary trace element for both plants and animals and is also an important resource. Nowadays, molybdenum is widely used in China's

semiconductor and optoelectronic industries because of its high melting point, high temperature resistance, good corrosion resistance and good thermal conductivity [1]. In addition, molybdenum is found in corrosion inhibitors, fertilizers and alloys [2]. Lianshan District of Huludao City is an important molybdenum industry base in China, accounting for one-third of the country's molybdenum resources. Many molybdenum industry

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enterprises in mining wastewater discharge treatment is not up to standard, after simple precipitation treatment directly into the water body, resulting in a serious molybdenum content exceeding the standard. As the exploration of molybdenum mines increases, the molybdenum pollution problem is becoming even more serious. Moreover, molybdenum pollution has been reported worldwide, such as in Erdenet Mine, Mongolia [3]; Wujintang Reservoir, China [4]; Modonkul River, Russian [5]; and Dabaoshan Mine, China [6]. For adults and older children, the recommended daily intake of molybdenum is 75-250 g [7]. The pollution of molybdenum to natural water environment will lead to the accumulation of molybdenum in food chain and cause chronic poisoning, so it is very important to remove heavy metal molybdenum from water environment. Excessive molybdenum intake can lead to adverse effects such as anemia, growth retardation, kidney stones and urethral calculi [7]. However, according to the effluent data of the mining wastewater from a molybdenum mining enterprise in Liaoning Province, the Mo (VI) concentration can reach about 20 mg/L. The World Health Organization (WHO) recommend the stheindicator for molybdenum in drinking water is 0.07 mg/L, the lack of feasible methods for Mo removal from aqueous solutions has become a growing concern for the sustainable use of natural resources [8].

Currently, many methods, e.g., chemical precipitation, ion-exchange and adsorption have been developed for Mo (VI) pollution control in aqueous solutions, and most of them have been demonstrated to be effective under certain environmental conditions [9-11]. Adsorption is now considered as an efficient and rapid way to remove heavy metals from water. Many researchers have investigated low-cost, biodegradable substitutes made from natural resources to remove toxic ions from wastewater [12-13]. Therefore, developing a low-cost and environmentally-friendly adsorbent with higher adsorption capacity is of importance and urgency [14]. Nanomaterials are widely used in the adsorption of organic dyes and heavy metals in water environment because of their small particle size, large surface area and high adsorbability [15-19]. Finding a nanoadsorbent that can quickly and effectively remove Mo (VI) in aqueous environment is a key issue. Chitosan is a natural polymer material, and a large number of active groups such as amino and hydroxyl groups exist on chitosan, which not only has good adsorption performance, but also has the characteristics of low cost and biodegradability [20]. Through the researches of Budnyak et al. [21], Brion-Roby et al. [22], Bertoni et al. [23], chitosan has strong adsorption on Mo (VI). The adsorption effect is good, at the same time, there is still a difficult problem of solid-liquid separation. Magnetic nanoparticles are easily magnetic separation and highly reusable adsorption materials. A large number of literatures have focused on the studies of magnetic nano-adsorption materials removing heavy metals such

as Cd, Cu, Sb, Cr [24- 27], while there are few studies on Mo. Fe_3O_4 is widely used in magnetic nanomaterials, so that the composite adsorption materials not only have high adsorption characteristics of nanoparticles, but also are endowed with magnetism. Hence, we consider coating chitosan with good adsorption performance on the Fe_3O_4 surface to form a composite adsorption material which is easy to magnetic separation.

Since there is no report on the improvement of Mo (VI) adsorption efficiency by chitosan-modified magnetic Fe_3O_4 nanoparticles, the adsorption characteristics of Fe_3O_4 /CTS magnetic nanoparticles by fitting the adsorption isotherms and adsorption kinetics was studied in this work. The chitosan-modified magnetic nanoparticle was prepared in this paper. The structure, morphology, magnetic properties and surface charge properties were characterized by XRD, IR, SEM and Zeta potential, respectively. In this paper, we determine the experimental conditions for adsorption Mo (VI) of composite adsorption materials and study the adsorption characteristics of Fe_3O_4 /CTS magnetic nanoparticles. Because the adsorption material is endowed with magnetism, the relationship between the regeneration times of the adsorption material and the removal rate of Mo (VI) is tested in this paper to study the adsorption and regeneration ability of Fe_3O_4 /CTS magnetic nanoparticles.

Materials and Methods

Materials

All chemical reagents were of analytical grade and the solutions in this study were prepared by deionized water. Chitosan, $\text{Na}_2\text{MoO}_4 \cdot 2\text{H}_2\text{O}$, ascorbic acid were purchased from Shanghai Lanji Technology Development Co., Ltd. NH_4SCN , $(\text{NH}_4)_2\text{Fe}(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$, $\text{Fe}(\text{SO}_4)_2$, NaOH , NaHCO_3 were purchased from Tianjin Hengxing Chemical Reagent Co., Ltd. Ethanol anhydrous, CH_3COOH , Paraffin liquid, Petroleum ether were purchased from Tianjin Fuyu Fine Chemical Co., Ltd. Glutaraldehyde, Propanone, $\text{NH}_3 \cdot \text{H}_2\text{O}$, H_2O_2 , H_2SO_4 , HCl were purchased from Tianjin Ruijin Special Chemicals Co., Ltd. Span-80 was purchased from Tianjin Damao Chemical Reagent Factory. PEG-20000 was purchased from Tianjin Institute of Guangfu Fine Chemical Industry.

Synthesis of Fe_3O_4 /CTS Magnetic Nanoparticles

Chitosan-modified magnetic nanoparticles were synthesized by H_2O_2 oxidation hydrothermal method. Dissolving 0.60 g of chitosan in 5% acetic acid, and then Fe_3O_4 magnetic nanoparticles were dissolved in chitosan solution by 1:1 mass ratio. Add 40 mL paraffin liquid, 1.5 mL Span-80, disperse in the Ultrasonic Cleaner (KQ-100E, China) for 20 min, and add 9 mL of 25% glutaraldehyde. After 4 h mechanical stirring,

the prepared substances were fully washed by petroleum ether, acetone and deionized water. Dry the resulting solid material in a Circulating Water Vacuum Pump (SHZ-D, Germany) at 60°C and grind it fully into a solid powder reserve [28].

Characterization of Magnetic Nanoparticles

The dry powder-like Fe₃O₄ magnetic nanoparticles and the chitosan-modified magnetic nanoparticles were identified by X-ray diffractometer. The scanning angle of the two particles was 10°-80°. Within the wavenumber range of 400 cm⁻¹ to 4000 cm⁻¹, three kinds of particles: Fe₃O₄ magnetic nanoparticles, chitosan and chitosan-modified magnetic nanoparticle were scanned by infrared spectrometer. The sample was dispersed in ethanol solution and observed by scanning electron microscope under the condition of 5 KV. Zeta potential measurements were formed by Malvern Nano-size Potentiometer at different pHs. The magnetic field intensity was measured when the magnetic field was in the range of -20000 Oe–20000 Oe.

Adsorption and Desorption Studies

Adsorption test: first, the pH of raw water was adjusted to 5.0 by 0.1 mol/L HCl and 0.1 mol/L NaOH. Weighing 0.40 g Fe₃O₄/CTS magnetic nanoparticles and mixing with 100 mL Mo (VI) wastewater in a conical flask. In the oscillating incubator (HZQ-X100, China), 150 r/min was used for oscillating adsorption, and then the magnetic separation of solid and liquid was obtained [29]. Mo (VI) concentration was determined by thiocyanate spectrophotometry [30], while adsorption capacity of Mo (VI) by Fe₃O₄/CTS magnetic nanoparticles was calculated by Eq. (1):

$$Q_t = \frac{(C_0 - C_t) \times V}{M} \quad (1)$$

...where Q_t (mg/g) was the adsorption capacity at t moment. C₀ (mg/L) and C_t (mg/L) was initial concentration of Mo (VI) and concentration after adsorption t time, respectively. V (mL) was the volume of water containing molybdenum solution, and M (g) was the adsorbent quality.

Regeneration test: using Fe₃O₄/CTS magnetic nanoparticles to adsorb Mo (VI) at room temperature. After reaching the adsorption equilibrium, the adsorption material was left in the conical bottle by the method of solid-liquid magnetic separation. Equipped with 0.01 mol/L NaOH solution, Fe₃O₄/CTS magnetic nanoparticles were ultrasonically desorbed by 20 mL NaOH solution for 30 min, then treated with deionized water ultrasonic for 30 min after magnetic separation. The regenerated magnetic nanoparticles were obtained after solid-liquid magnetic separation. The reusability characteristics of magnetic adsorption materials were

investigated. The adsorption tests were carried out at room temperature [31].

Results and Discussion

Characterization of Fe₃O₄/CTS Magnetic Nanoparticles

X-ray diffractometer scans of a) Fe₃O₄ and b) Fe₃O₄/CTS were shown in Fig. 1. In the XRD patterns of the two adsorbed materials, there were six typical diffraction peaks around 30.3°, 35.7°, 43.3°, 53.4°, 57.2° and 62.8°, corresponding to different crystal planes (220), (311), (400), (422), (511) and (440), respectively. Comparing the curve of Fe₃O₄/CTS (Fig. 1b) with the standard diffraction card Fe₃O₄ (Fig. 1a), there were no other crystal peaks, indicating that pure Fe₃O₄ magnetic nanoparticles were prepared by H₂O₂ oxidation method. Whereas the chitosan-modified magnetic nanoparticles were completely consistent with the diffraction peaks of Fe₃O₄ magnetic nanoparticles, indicating that the phase transition of Fe₃O₄ nanoparticles was not caused during the binding of chitosan, and the crystal structure did not change [32-34].

FTIR spectra of a) CTS, b) Fe₃O₄, and c) Fe₃O₄/CTS were illustrated in Fig. 2. It was shown that the stretching vibration of active groups such as -NH₂ and -OH exist in the FTIR spectra of CTS (Fig. 2a), so there were characteristic absorption peaks near 3360.9 cm⁻¹ in FTIR spectra. After coordination of Fe₃O₄ (Fig. 2b), the characteristic absorption peak of Fe₃O₄/CTS (Fig. 2c) magnetic nanoparticles was obviously offset, and the absorption peak moved to 3330.9 cm⁻¹ in the direction of low wavenumber. This might be due to the weakening of N-H stretching vibration on chitosan after the Fe₃O₄ composite, and the formation of N-Fe bond enhances. The C-H bending

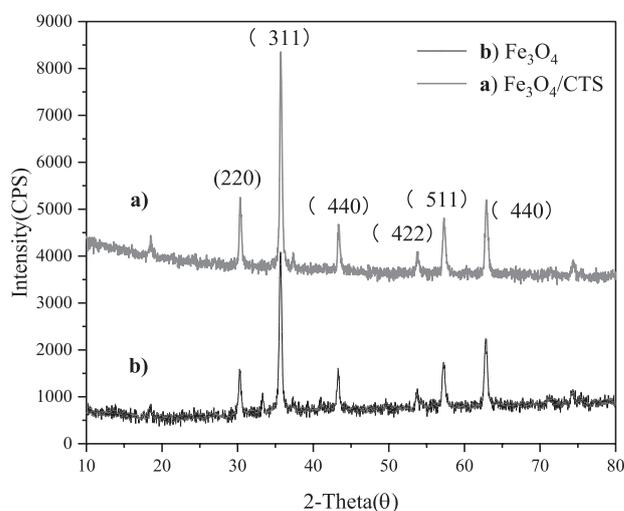


Fig. 1. X-ray diffractometer scans of a) Fe₃O₄ and b) Fe₃O₄/CTS.

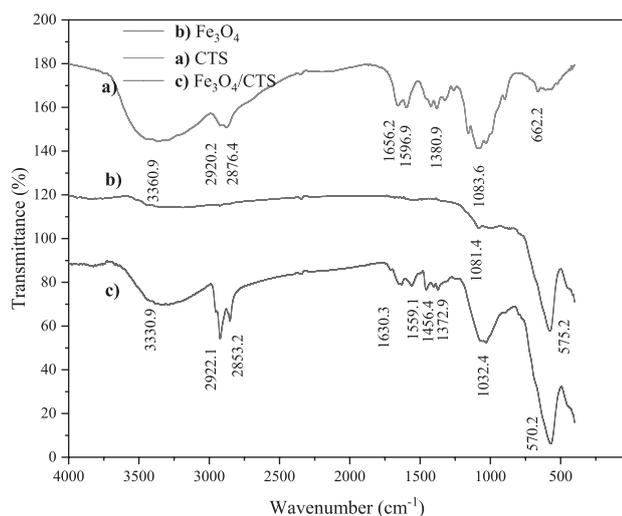


Fig. 2. FTIR spectra of a) CTS, b) Fe_3O_4 , and c) $\text{Fe}_3\text{O}_4/\text{CTS}$.

vibration peak at 1380.9 cm^{-1} also shifted to the low wavenumber due to the formation of composites. In addition, due to the expansion vibration of $\text{C}=\text{N}$, there was a characteristic absorption peak at 1630.3 cm^{-1} , indicating that the crosslinking agent glutaraldehyde had an effect on NH_2 . $\text{Fe}-\text{O}$ was the characteristic peak of Fe_3O_4 , which appeared at 575.2 cm^{-1} of Fe_3O_4 nanoparticles and 570.2 cm^{-1} of $\text{Fe}_3\text{O}_4/\text{CTS}$

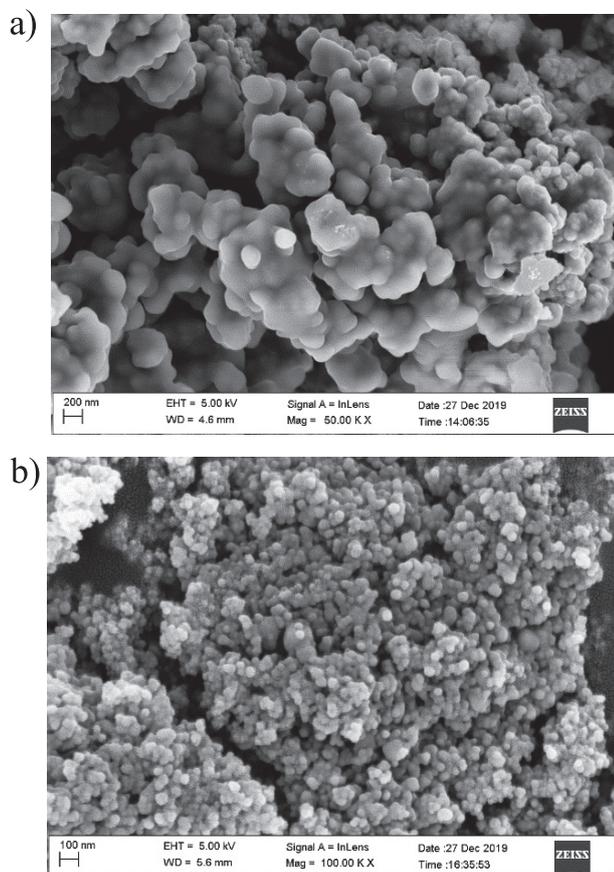


Fig. 3. SEM images of a) Fe_3O_4 and b) $\text{Fe}_3\text{O}_4/\text{CTS}$.

nanoparticles, respectively, indicating that $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles were formed successfully after the chemical effect of Fe_3O_4 on chitosan.

SEM images of a) Fe_3O_4 and b) $\text{Fe}_3\text{O}_4/\text{CTS}$ were shown in Fig. 3. Both of them were spherically grained. However, the agglomeration of Fe_3O_4 nanoparticles still occurred after encapsulated by CTS. The particles indicated that agglomeration might be formed in the drying process. Because the particle size was very small, the increase of specific surface area led to the increase of active sites on the surface of the particles. The chemical activity increased, that was, chemical stability was poor, which led to the phenomenon of agglomeration.

Fig. 4 showed the hysteresis curves of two magnetic nanoparticles measured at the temperature of 300 K. The saturation magnetization of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles (Fig. 4a) was 51.7 emu/g , which decreased compared with that of Fe_3O_4 magnetic nanoparticles (Fig. 4b) with a saturation magnetization of 87.1 emu/g . The reason was that chitosan was successfully coated on the surface of Fe_3O_4 particles and weakened its magnetization. It could be seen from the hysteresis curves that the coercivity and remanence of the two magnetic nanomaterials were close to zero, which indicated that both of them had superpara magnetic properties. The magnetic properties of composite nanomaterials could be easily separated from the liquid by external magnetic field after the adsorption reaction was completed. Magnetic composites were easy to recycle and reuse.

The zeta potentials of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles under different pHs were shown in Fig. 5, which determined the surface charge change of the composite when suspended in aqueous medium. Zero potential point (pH_{pzc}) of nanocomposites was observed near $\text{pH} = 5.3$. When $\text{pH} < \text{pH}_{\text{pzc}}$, the surface of $\text{Fe}_3\text{O}_4/\text{CTS}$ nanoparticles was positively charged, and it was easy to adsorb anions by electrostatic

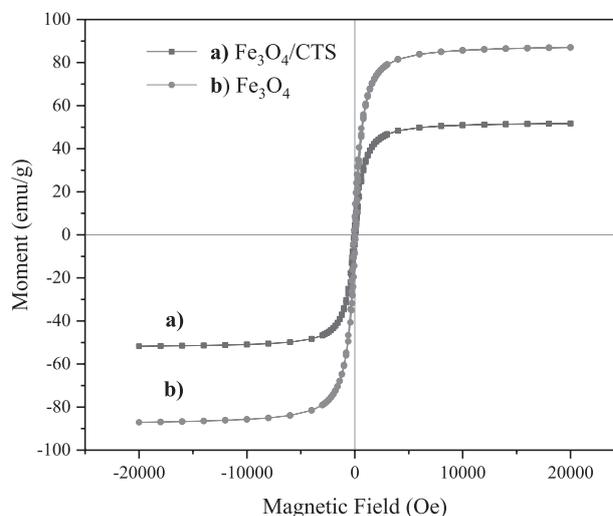


Fig. 4. Hysteresis curves of a) Fe_3O_4 and b) $\text{Fe}_3\text{O}_4/\text{CTS}$.

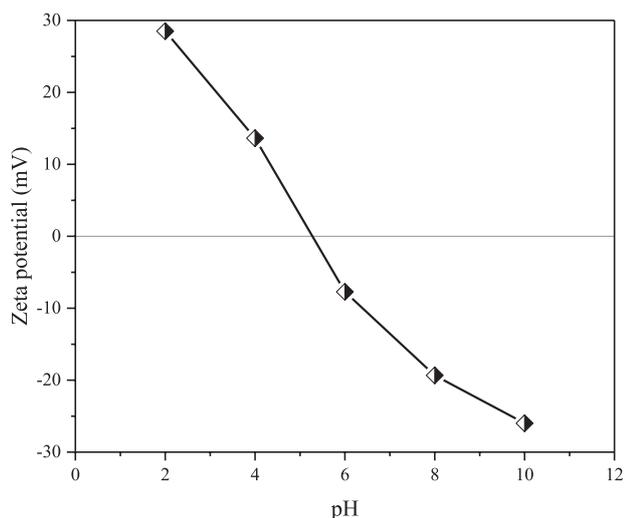


Fig. 5. Zeta potential of Fe₃O₄/CTS.

attraction. However, when $\text{pH} > \text{pH}_{\text{pzc}}$, the surface of the nanoparticles was negatively charged, and the adsorption effect of anion by electrostatic repulsion would be reduced, which tended to adsorb cations [35].

Effect of Initial pH on Adsorption

The removal of Mo (VI) by nanocomposites under different pH conditions was studied by adsorption of 0.40 g Fe₃O₄/CTS magnetic nanoparticles in the pH range of 2-10. Removal rates of Mo (VI) under different pH conditions were shown in Fig. 6. In Fig. 6, the initial pH of the solution had a great effect on the removal of Mo (VI). The removal rate of Mo (VI) by Fe₃O₄/CTS magnetic nanoparticles increased first and then decreased with the increase of pH value. The highest removal rate of Mo (VI) was 94.35% when the

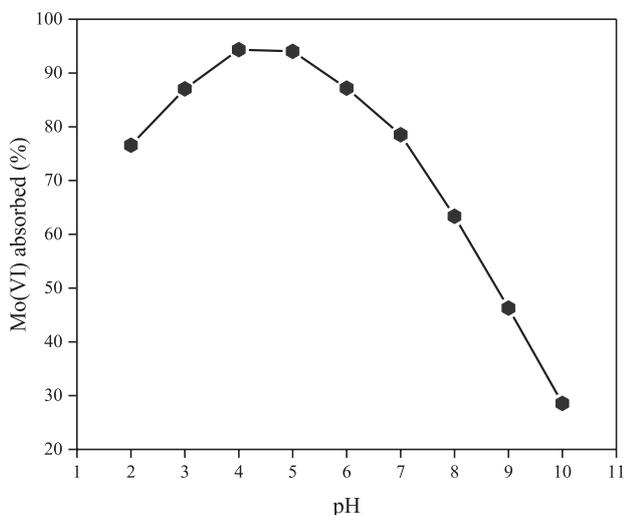


Fig. 6. Removal rates of Mo (VI) under different pHs. Condition: T = 298 K, solution volume 100 mL, nano-magnetic Fe₃O₄/CTS 0.40 g, 100 mL of 30 mg/L Mo (VI), agitation time of 30 min.

initial pH value was 4, and the maximum adsorption capacity of magnetic composites was 7.08 mg/g. The presence of Mo (VI) in water and the surface charge of Fe₃O₄/CTS magnetic nanoparticles were the key factors for the adsorption of Mo (VI) [36]. Guibal et al. [36] studied the presence of molybdenum at concentrations of 10 mg/L and 100 mg/L in different pH values. In neutral to weak acidity, it mainly existed in the form of MoO₄²⁻. The oxygen anions of HMoO₄⁻, H₂MoO₄, Mo₇O₂₁(OH)₃³⁻ and Mo₇O₂₄⁶⁻ were also present under acidic conditions. All Mo (VI) except H₂MoO₄ were in the form of oxygen anion and Mo₁₉O₅₉⁴⁻ only existed in high concentrations of Mo (VI) solution. Therefore, the concentration of Mo (VI) might also be a factor affecting adsorption efficiency. When initial pH was less than 4, adsorption of Mo (VI) decreased, probably because a stronger the acidity results in a greater proportion of H₂MoO₄. The pH value did not reach the equipotential point (pH_{pzc}). The surface of Fe₃O₄/CTS magnetic nanoparticles was positively charged, and it easily adsorbed the oxygen anion of molybdenum. The presence of H₂MoO₄ might be the main reason for the decrease of Mo (VI) removal rate. Although adsorption efficiency decreased in this range, the removal rate was over 75%. At pH = 4, H₂MoO₄ was almost absent from the solution, and its adsorption efficiency was the highest. As the pH continued to increase, the surface of the Fe₃O₄/CTS magnetic nanoparticles became even more negatively charged, and the generation of electrostatic repulsion led to a significant decrease in Mo (VI) removal rate.

Adsorption Kinetic Models

Adsorption kinetics was very important in evaluating the properties of adsorption materials and studying the potential mechanism of the adsorption process [37]. At room temperature, the pH of molybdenum-containing wastewater was adjusted to 4, 100 mL water sample was added to the conical bottle and 0.04 g Fe₃O₄/CTS magnetic nanoparticles were used for oscillating adsorption at 150 r/min speed until adsorption equilibrium was reached. The adsorption process of Fe₃O₄/CTS magnetic nanoparticles on Mo (VI) was fitted by pseudo-first-order kinetic equation [38], pseudo-second-order kinetic equation [39] and particle diffusion model respectively, and their equations were expressed as (2)-(4):

$$\ln(Q_e - Q_t) = \ln Q_e - \frac{k_1}{2.303} t \quad (2)$$

$$\frac{t}{Q_t} = \frac{1}{k_2 Q_e^2} + \frac{t}{Q_e} \quad (3)$$

$$Q_t = k_p t^{1/2} + C \quad (4)$$

Table 1. Comparison of the two kinetic models for Mo (VI) adsorption on Fe₃O₄/CTS magnetic nanoparticles at 298 K.

Kinetic model		Mo (VI) ion
Mo (VI) mass concentration(mg/L)		30
Q _{c,exp} (mg/g)		7.46
Pseudo-first-order		
k ₁ (min ⁻¹)		0.036
Q _c (mg/g)		0.803
R ²		0.7319
Pseudo-second-order		
k ₂ (g/(mg·min))		0.064
Q _c (mg/g)		7.47
R ²		0.9999
Internal diffusion model		
k _{i,1}		0.345
k _{i,2}		0.139
k _{i,3}		0.014

...where t (min) was the contact time, Q_c (mg/g) and Q_t (mg/g) was the adsorption capacity of Fe₃O₄/CTS magnetic nanoparticles on Mo (VI) at adsorption equilibrium and at adsorption of t min, respectively. C was boundary layer constant. k_1 and k_2 was the velocity constants of pseudo-first-order and pseudo-second-order dynamics, respectively. k_p was particle diffusion dynamics parameter. When the adsorption rate was proportional to the number of adsorption sites unoccupied by the adsorbent, the adsorption process conformed to pseudo-first-order kinetics. If the occupancy rate of the binding site was proportional to the square of the unoccupied site, adsorption conformed to pseudo-second-order kinetics [40].

For sodium molybdate solution with an initial concentration of 30 mg/L, the adsorption rate of Mo (VI) was very fast within 30 min and reached adsorption equilibrium for about 60 min. Per gram of Fe₃O₄/CTS magnetic nanoparticles were able to adsorb 7.4 mg Mo (VI), and the concentration of Mo (VI) was 0.38 mg/L at this point. Adsorption kinetics of Mo (VI) were shown in Fig. 7 and the fitting results were shown in Table 1. From Fig. 7 and Table 1, Q_t/t had a good linear relationship with t , $R^2 = 0.9999$, the saturated adsorption capacity was similar to the experimental results, and the pseudo-second-order kinetics could better fit the adsorption of Mo (VI) by Fe₃O₄/CTS magnetic nanoparticles. This suggested that the steps to control the adsorption rate might be limited by the chemical adsorption of shared or exchanged electrons between the adsorbent and adsorbate [14]. The driving force of this process depended on the square of the adsorption vacancy point [35], which could explain the

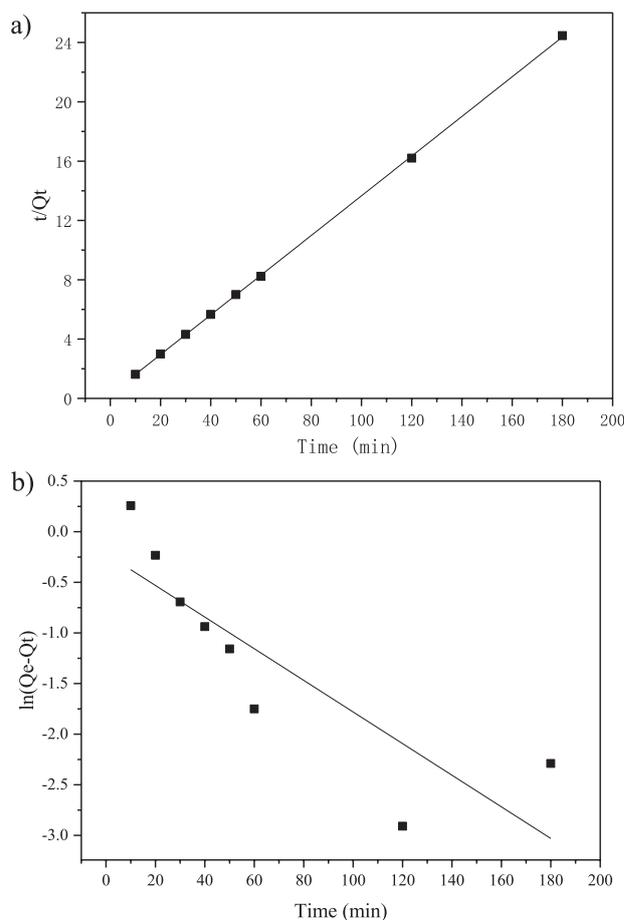


Fig. 7. Adsorption kinetics of Mo (VI). (a) pseudo-second-order kinetic model and (b) pseudo-second-order kinetic model. Condition: $T = 298$ K, $pH = 4.0$, Fe₃O₄/CTS magnetic nanoparticles 0.40 g, 100 mL of 30 mg/L Mo (VI).

fast adsorption rate in the early stage of adsorption, but in the later stage of adsorption, the decrease of adsorption vacancy points on Fe₃O₄/CTS magnetic nanoparticles and the decrease of chemical driving force brought by the decrease of Mo (VI) concentration led to slower adsorption rates [41].

Intramolecular diffusion model took $t^{0.5}$ as the abscissa and adsorption capacity as the ordinate. The fitting results were shown in Fig. 8. The straight line in the diagram does not pass through the origin, indicating that the internal diffusion of particles is not the only step to control the adsorption rate. Therefore, the adsorption of Mo (VI) by Fe₃O₄/CTS magnetic nanoparticles may be divided into three stages. According to the fitting data of Table 1, with the increase of adsorption time, the adsorption rate becomes slower. The slope of the three-segment broken line was the diffusion rate constant of particle diffusion: $k_{i,1} > k_{i,2} > k_{i,3}$. The adsorption process of Mo (VI) on Fe₃O₄/CTS magnetic nanoparticles may be that in the first stage of adsorption Mo (VI) bind to the adsorbable active sites of functional groups on the surface of Fe₃O₄/CTS magnetic nanoparticles, and the initial

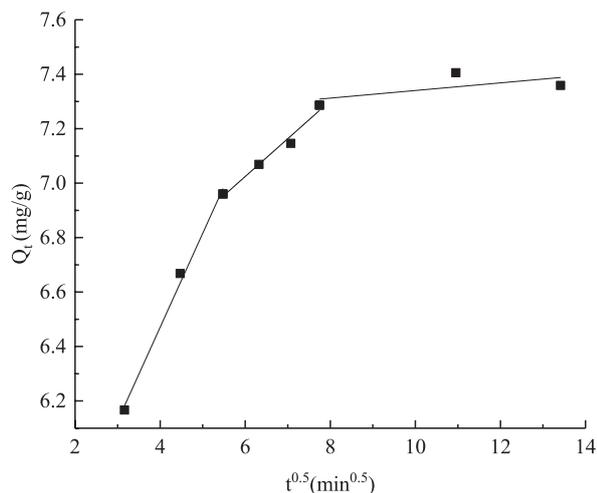


Fig. 8. Intra particle diffusion model of Mo (VI) adsorbed on $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles. Condition: $T = 298 \text{ K}$, $\text{pH} = 4$, $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles 0.40 g , 100 mL of 30 mg/L Mo (VI).

adsorption rate was fast. When the number of active sites on the surface decreases gradually, it began to enter the second stage of adsorption. The adsorption sites on the surface of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles were basically occupied by Mo (VI). While Mo (VI) began to enter the micropores of the magnetic nanoparticles and were adsorbed by the pores. At the third stage of adsorption, the adsorption sites on the surface of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles and in the pore basically saturated. The adsorption reached equilibrium state at this time. Through the model of particle diffusion, it could be concluded that the adsorption of Mo (VI) on $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles is not only chemical chelation, but also internal diffusion.

Adsorption Isotherms

Adsorption isotherms studied the relationship between adsorption capacity and Mo (VI) concentration of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles at constant temperature. The experiments were carried out at room temperature. The concentration in the range of $30\text{--}210 \text{ mg/L}$ was selected as the initial concentration of Mo (VI) using Langmuir model [42] and Freundlich model [43] fitting of adsorption process. Equations were expressed as (5)-(6):

$$\frac{C_e}{Q_e} = \frac{C_e}{Q_m} + \frac{1}{K_L Q_m} \quad (5)$$

$$\ln Q_e = \ln K_F + \frac{1}{n} \ln C_e \quad (6)$$

...where Q_m (mg/g) was the largest monolayer adsorption capacity, Q_e (mg/g) indicated adsorption capacity

Table 2. Parameters of Langmuir and Freundlich isotherm equations for the adsorption of Mo (VI) at 298 K .

Adsorption model	Mo(VI) ion
Langmuir isotherm	
Q_m (mg/g)	35.54
K_L (L/mg)	0.08
R^2	0.9978
Freundlich isotherm	
n	2.36
K_F ($\text{mg}^{1-1/n}\text{L}^{1/n}/\text{g}$)	5.42
R^2	0.9501

at adsorption equilibrium, C_e (mg/L) indicated the concentration of Mo (VI) in adsorption equilibrium, K_L and K_F was Langmuir and Freundlich constants, respectively, $1/n$ was the adsorption index.

Table 2 showed the fitting results. The results showed that although the adsorption equilibrium concentration C_e of Mo (VI) increased with the increase of initial concentration, the equilibrium adsorption capacity of the adsorbed material Q_e was also increasing. Using linear regression to process the isotherm data, Fig. 9a) and 9b) indicated the linear relation of C_e/Q_e to C_e and $\ln Q_e$ to $\ln C_e$. In Fig. 9a), Q_m and K_L could be calculated and the correlation coefficient $R^2 = 0.99777$ was closer to 1 than that of Freundlich model. Freundlich model was more suitable for adsorption on highly hybrid surfaces lacking saturated adsorption platforms [43], while Langmuir model was more suitable for adsorption of Mo (VI) by $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles at room temperature. The average saturated adsorption capacity of 31.46 mg/g was very close to the calculated maximum adsorption capacity $Q_m = 35.54 \text{ mg/g}$. Langmuir model assumed that the adsorption effect was monolayer, so $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles were monolayer adsorption. Each active site could only adsorb one molecule, and all the adsorption sites were in the same priority, forming a uniform monolayer adsorption membrane [29]. Table 3 compared the $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles with other adsorbed materials in the removal of Mo (VI). Obviously, the removal performance of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles to Mo (VI) was good, the adsorption capacity was 35.54 mg/g , which was better than that of D2EHFA-functionalized magnetite (25.8 mg/g) [44], zeolite-supported magnetite (18.02 mg/g) [45], sulfuric acid-modified cinder (31.08 mg/g) [46]. Even though there were adsorbents like polysiloxane modified with DETAP (111.84 mg/g) [47] and QA-MNP (65 mg/g) [35] with excellent adsorption properties, the preparation of these adsorbents might be complex and costly, which was not conducive to production and application.

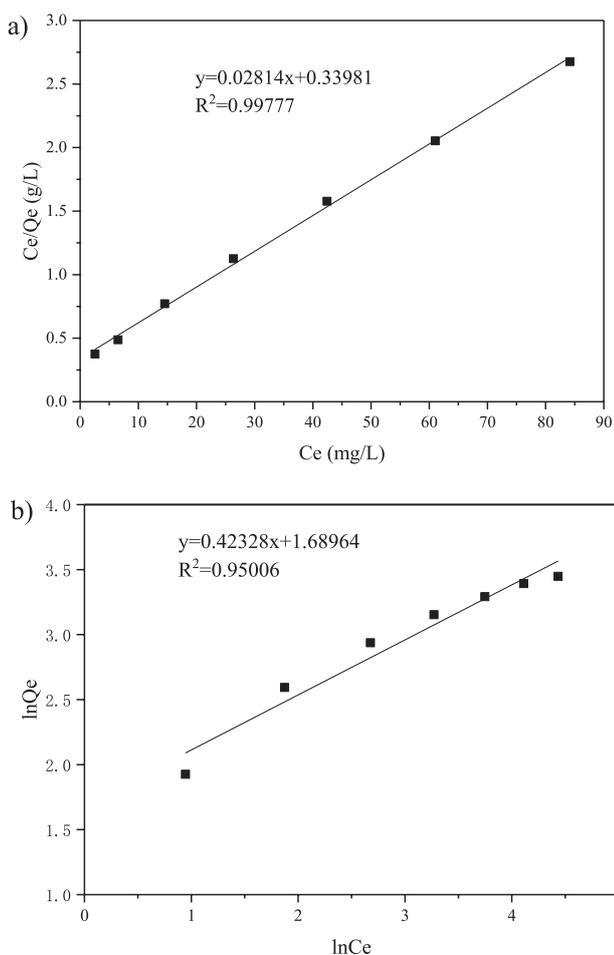


Fig. 9. Adsorption isotherms of Mo (VI). a) Langmuir model and b) Freundlich model. Condition: $T = 298$ K, $\text{pH} = 4$, nano-magnetic $\text{Fe}_3\text{O}_4/\text{CTS}$ 0.40 g, Mo initial concentration 30-210 mg/L, solution volume 100 mL.

Desorption and Regeneration Test

The reusability of adsorbed materials was an important factor for measuring its efficiency. In order to make the adsorption process of Mo (VI) by $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles more economical and effective, the desorption experiment of this adsorption material was studied. Using 0.01 mol/L NaOH solution for ultrasonic treatment of adsorbed equilibrium water

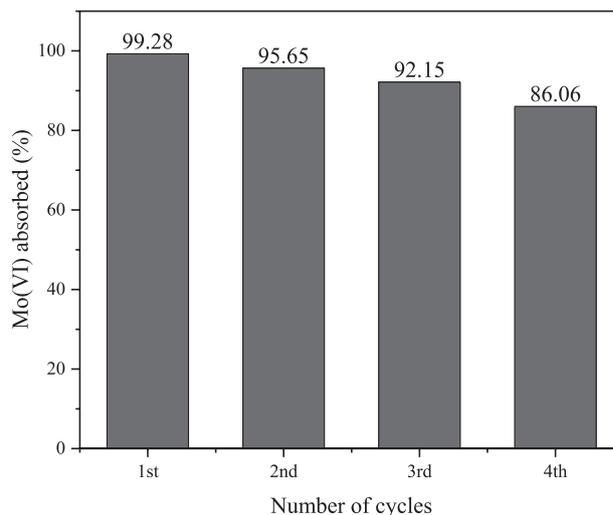


Fig. 10. Performance of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles by multiple cycles of regeneration. Desorption condition: 0.01 N NaOH solutions for 30 min at 298 K.

samples, the magnetic separation of solid solution was used to obtain the regenerated composite nanomaterials, and the relationship between the regeneration times of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles and the removal rate of Mo (VI) were studied. The performance of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles by multiple cycles of regeneration were shown in Fig. 10. The results showed that the removal rate of Mo (VI) was 99.28% at adsorption equilibrium, and the adsorption capacity decreased gradually with the increase of regeneration times. The removal rate of Mo (VI) after one regeneration of $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles was 95.65%, and the removal rate decreased to 92.15% after two regenerations, which still reached more than 90%. According to the similar consistency principle and electrostatic repulsion [13], OH^- on NaOH solution could replace Mo (VI) at the adsorption site of $\text{Fe}_3\text{O}_4/\text{CTS}$ particle, which enabled the regenerated adsorbate material to be reused and obtain the ability to adsorb Mo (VI) again [48]. However, the reason for the decrease of adsorption capacity might be that there was still a small amount of Mo (VI) on the regenerated $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles, leading to the

Table 3. Comparison between adsorption of Mo (VI) by $\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles and other adsorbents.

Adsorbents	Adsorption capacity (mg/g)	References
Polysiloxane modified with DETAP	111.84	Sharma et al. (2012)
Zeolite-supported magnetite	18.02	Verbinnen et al. (2013)
D2EHPA- functionalized magnetite	25.8	Agarwal et al. (2016)
$\text{Fe}_3\text{O}_4@\text{SiO}_2$	31.08	Hassanpour et al. (2016)
QA-MNP	65	Gaete et al. (2019)
$\text{Fe}_3\text{O}_4/\text{CTS}$ magnetic nanoparticles	35.54	This study

decrease of adsorption vacancy as well as to a removal rate that was lower than the initial adsorption rate [29, 41].

Conclusions

Fe₃O₄/CTS magnetic nanoparticles were synthesized by reverse suspension cross-linking. The chitosan-modified magnetic nanoparticles were completely consistent with the diffraction peaks of Fe₃O₄ magnetic nanoparticles, indicating that the phase transition of Fe₃O₄ nanoparticles was not caused during the binding of chitosan, and the crystal structure did not change. FTIR spectra showed that Fe₃O₄/CTS magnetic nanoparticles were formed successfully after the chemical effect of Fe₃O₄ on chitosan. Zero potential point (pH_{pzc}) of nanocomposites was observed near pH = 5.3. The results showed that the removal effect of molybdenum was the best under the condition of pH = 4.0, and the removal rate is 94.35%. The adsorption behavior of Mo (VI) on Fe₃O₄/CTS magnetic nanoparticles followed the Langmuir isotherm, the maximum adsorption capacity was 35.54 mg/g. Pseudo-second-order kinetics can accurately describe the adsorption process. At the same time, the adsorption process of molybdenum on Fe₃O₄/CTS magnetic nanoparticles had the effect of intraparticle diffusion. Through the experimental study of desorption and regeneration of nano-adsorbed materials, the removal rate of molybdenum could still reach 92.15% after two regeneration, indicating that Fe₃O₄/CTS magnetic nanoparticles had a good reproducibility, strong reusability and good removal effect of Mo (VI). The Fe₃O₄/CTS magnetic nanoparticles were a kind of good adsorption material with great prospects for rapid and efficient adsorption, which could be reused and had a broad research prospect.

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Conflict of Interest

The authors declare no conflict of interest.

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