

Advanced Decomposition of Coking Wastewater in Relation to Total Organic Carbon Using an Electrochemical System

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Abstract

This research included experiments on the advanced treatment of coking wastewater by electrochemical reactor. The results showed that a favorable operating condition could be obtained through the electrochemical course, when electrolysis time was 60 min, current density was 8 A, and electrode span was 1 cm. Under this condition, total organic carbon (TOC) removal efficiency could reach 73%, and the removal mechanism had also been analyzed. In addition, the electrodes' morphology characteristics and components were analyzed through a scanning electron microscope, and the results showed that on the surface of anode and cathode electrodes there were quite a few cracks, the amount of which increased after use. Meanwhile, the content of chemical elements on the electrodes' surface changed obviously after use – especially Ti content (from 39.66% to 92.69%) on the anode and Fe content (from 72.57% to 53.66%) on the cathode. The result was probably caused by the redox reaction, namely the shedding of the coating on the anode electrode's surface, and reactions around the cathode electrode's surface.

Keywords: remote sensing, satellite images, lakes, oligotrophy

Introduction

As a new type of electrochemical processing, electrochemical technology has recently been paid more and more attention in the wastewater treatment field due to its advantages of larger specific surface area, better transfer effect, higher current efficiency, and space-time yield [1-3]. Electrochemical treatment of wastewater is a

relatively new type of wastewater treatment technology that is gradually developing into a promising technology [4-5]. Meanwhile, it has been successfully applied in the purification of wastewater from dye production [6] and from textile wastewater containing cyanides [7], phenols [8-9], and domestic sewage [10].

Coking wastewater is generated from coal coking, coal gas purification, and by-product recovery processes of coking [11-14]. It contains complex inorganic and organic pollutants such as ammonium, sulfate, cyanide, thiocyanate, phenolic compounds, polynuclear aromatic hydrocarbons, and polycyclic nitrogen-containing acyclic compounds,

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etc. [15-17]. Meanwhile, most academic studies have focused on the reaction mechanism and determining the reaction intermediate for the electrochemical method, and further studies on the influence of various factors for experimental effect also were deeply investigated [18-22].

Besides, in order to get better removal efficiency, RuO_2 and IrO_2 coated with titanium plates were used as anodes of the electrodes device in our experiment. Therefore, coking wastewater was applied in the single factor experiment, and the optimal operating parameters were carried out through the experimental course. Meanwhile, total oxygen carbon (TOC) was analyzed for treatment efficiency from coking wastewater, and effects of various factors could provide the basis for the popularization and application of this technology.

Experimental

Electrolytic System

The electrochemical device consisted of the electrolytic cell, electrode plates, and DC power supply. The electrolytic cell was made of synthetic glass, with a dimension of $140 \times 90 \times 130$ mm and an effective volume of 1.3 L. Tests of electrolysis were carried out using an electrolytic device consisting of stainless steel cathode plates and titanium anode plates coated with noble metals, including Ru, Ir, etc. The specification for each plate was $130 \times 65 \times 1$ mm. The schematic diagram of the main apparatus is shown in Fig. 1.

Influent Quality

Wastewater for this study was fed from a secondary sedimentation tank in a coking plant in Tangshan. The composition of the coking wastewater used in all experiments is shown in Table 1.

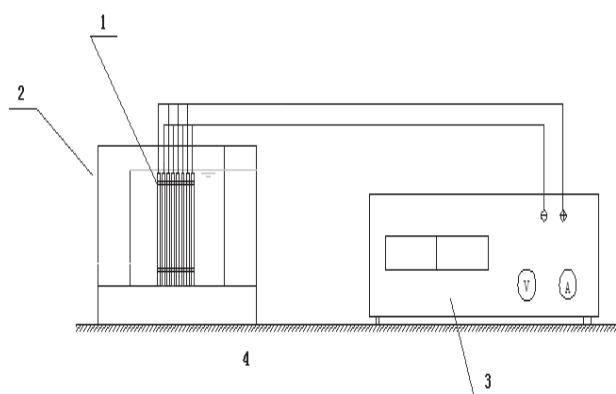


Fig. 1. Schematic diagram of the experimental apparatus: 1. electrode plates, 2. electrolytic cell, 3. electric power source, 4. experiment table.

Table 1. Raw water quality and recycled water indexes.

Parameter	TOC (mg/L)	Chroma (mg/L)	pH
Concentration	182.3~230.9	100.2~136.5	5.9~6.7

Results and Discussion

Influence of Electrolytic Time

As is shown in Fig. 2, the removal rate of TOC kept rapid growth with the length of electrolytic time in the first 60 min and became flat in the last 40 min. And organic compounds could be obviously degraded from the wastewater. However, there was little function to the treatment when prolonging reaction time after 60 min. Thus the optimal electrolysis time was 60 min.

Influence of Electrolytic Current

As seen in Fig. 3, the removal rates of TOC could reach 73% when electrolysis current was 8 A, while it increased quickly with the current density. However, the removal rates of TOC decreased if the current exceeded 8 A, and wastewater temperature rose quickly. Meanwhile, reaction driving-power and electric energy consumption increased with the TOC removal rate. But there were removal and energy input requirement extreme for pollutants in the electrochemical system. Besides, because the input energy achieved the extreme circumstance, there was a higher removal efficiency and better redox reaction with the input energy increasing. On the contrary, with the decreasing of input energy, adverse reactions happened and the current efficiency decreased instead of increasing removal efficiency. Therefore, the proper current of the experiment was 8A.

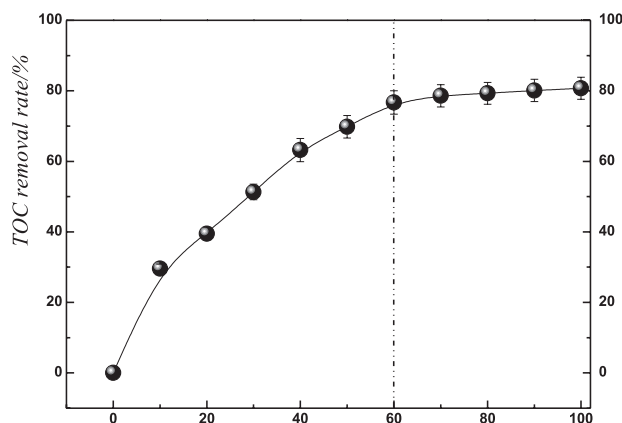


Fig. 2. Influence of electrolytic time on TOC removal rate (when current density is 8 A and electrode span 1 cm; percentages of TOC removal are shown as the means \pm s.d., $n = 3$).

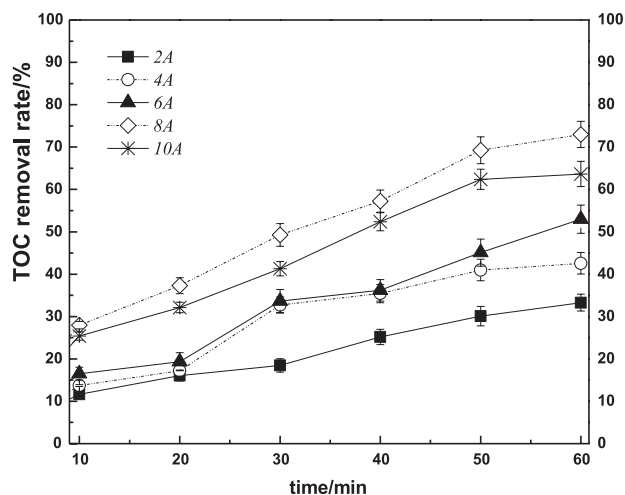


Fig. 3. The influence of electrolytic current on TOC removal rate (when electrolysis time was 60 min and plate spacing 1 cm; percentages of TOC removal are shown as the means \pm s.d., $n = 3$).

Influence of Electrode Span

As seen in Fig. 4, the treatment effect of TOC reached a maximum of about 67% when plate spacing was 1 cm, while the removal rates of TOC were obviously influenced by it. The smaller the plates spacing, the larger the current on the electrode. And the concentrations of H_2O_2 and $\bullet OH$ were higher, which influenced the degradation of organic matter. When the plate spacing become too small, the current on the electrode decreased rapidly as the generation of spark in the reaction process increased with electrode resistance, and then electrolysis was reduced. Meanwhile, when plate spacing increased, there would be higher system resistance that resulted in a decrease in removal rates. As an influence on the removal of TOC with

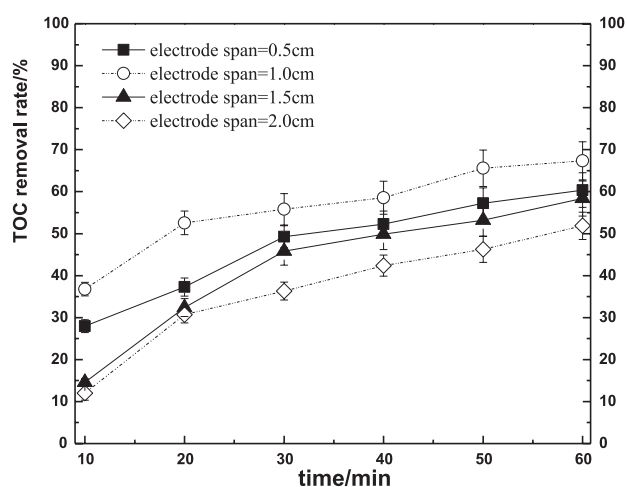


Fig. 4. Influence of electrode span on TOC removal rate (when electrolysis time is 60 min and current density 8 A; percentages of TOC removal are shown as the means \pm s.d., $n = 3$).

current density, the reasons came down to whether input energy was extreme energy. Besides, Fig. 4 indicated that the input power was approaching extreme power mostly when plate spacing was 1 cm.

Anode Morphology Characteristic and Component Analysis

As seen in Figs 5 and 6, electrode coating exhibited favorable physical characteristics before use, but some cracks still existed on its surface. After use, fair-sized depletions and coating cracks began to appear on its surface, which was the reason why electrode removal effectiveness decreased. Meanwhile, nascent oxygen was generated during electrolysis, but some cracks appeared on the active coatings, so part of the oxygen may penetrate into the titanium substrate surface by diffusion or migration patterns, and then generate a non-conductive TiO_2 with titanium and produce the reverse PN junction conductance, while electrolysis liquid through the cracks may also invade the titanium base. There would be a titanium oxide matrix to appear, which caused titanium substrate with an active coating interface and exfoliation corrosion. Meanwhile, this made electrode

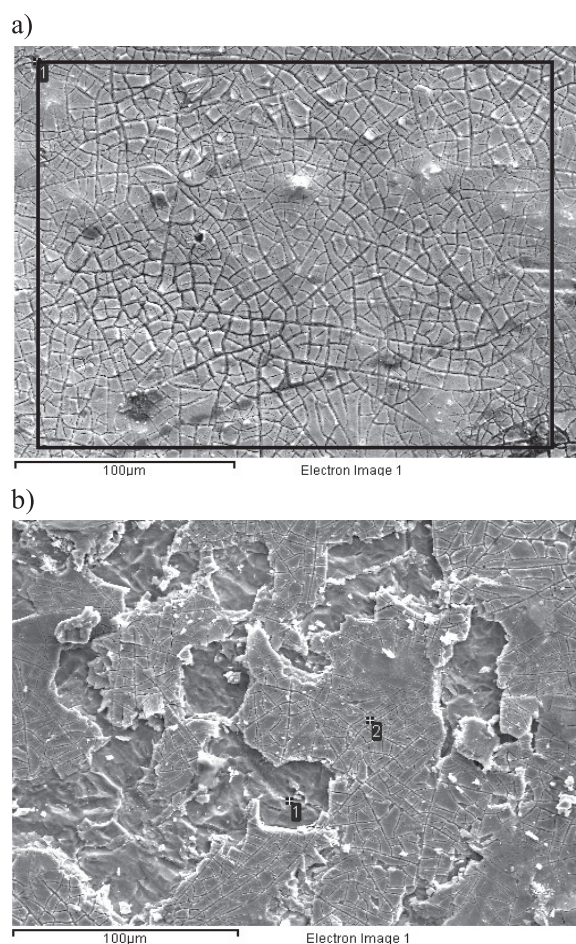


Fig. 5. $Ti/RuO_2 \cdot IrO_2$ morphology characterization of anodic coating: a) before use, b) after use.

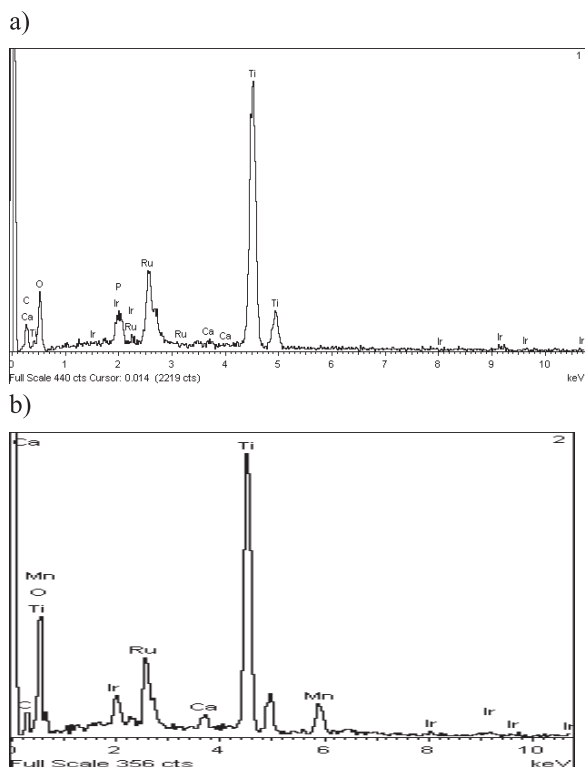
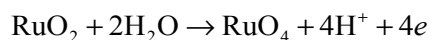


Fig. 6. Anode EDS maps: a) before use, b) after use.

potential increase, and electrode potential rose further and promoted the dissolution of the coating and titanium substrate oxidation; this severe circle would eventually lead to electrode failure.

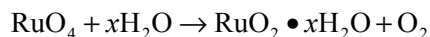
In addition, Ti content in surface coating could achieved 92.69% after use, which indicated that the coatings almost peeled off completely, and one point on the titanium-based surface in Fig. 5b) could approve it. Meanwhile, when the coatings peeled off to a certain degree, the passivation would not carry out. Besides, Ca content in coating increased with a small quantity, which indicated that no serious calcium anode fouling occurred on the electrode. However, higher Mn content appeared on used anode surface, and it was relatively high. Meanwhile, because there was some heavy metal content in the coal, the coking wastewater often contained considerable metal ions, and these metal ions formed organometallic complexes with phenols in wastewater and organic vehicles. These complexes were poisonous and deemed to be the “three-induced” effect on nature, which had important environmental significance. Meanwhile, heavy metals were found to have a plentiful presence in coking wastewater, and the average Mn content (10 samples) was about 3.0 mg/l at low temperature, compared with 1.75 mg/l at high temperature. Besides, Mn, as a medium, could play an important role in dielectric oxide, and Mn ions exhibited a beneficial effect for the presence of oxidation. However, Ru content decreased, indicating that Ru dissolution occurred, and when the anode potential was more positive than 1.387 V vs. SHE, then RuO_2 was oxidized to RuO_4 :



or:



RuO_2 might form RuO_4 in solution, and then RuO_4 could be decomposed:



As shown in above, the black precipitates could be considered $\text{RuO}_2 \cdot x\text{H}_2\text{O}$, which would be deposited on the electrode surface.

Under conditions of the alkaline environment, RuO_2 was oxidized to RuO_4 , but OH^- could discharge and produce oxygen under low potential, thereby inhibiting oxidation of catalytic activity in coatings. Compared with acid environment, the operating life of coatings would be long enough for 5~6 times. In addition, acid solution contained chloridion, as Ru and Cl^- would combine under different valence states, and soluble $\text{RuCl}^{(n-m)}$ (chloride complex ions) would be generated. Chloridion concentration was relatively high in coking wastewater, and anode existed oxygen evolution reaction, thus obvious dissolution occurring in anode coatings would be reasonable. Meanwhile, the electrode would be not worked when the anode surface catalysis coatings dissolved to an extent.

As seen in Tables 2 and 3, Ir content was minor in coatings, which induced the chlorine evolution potential and oxygen evolution potential that were not high, and the difference was small. Meanwhile, increasing Ir content could raise the oxygen evolution potential, which could enhance the chlorine evolution efficiency to lessen the coatings depletion.

Cathode Morphology Characteristic and Component Analysis

As seen in Fig. 7, crack morphology of the cathode surface could be in favor of hydrogen escape along with

Table 2. $\text{Ti/RuO}_2\text{-IrO}_2$ anode electrode coating composition (before use).

Element	Weight%	Atomic%
C K	6.43	14.34
O K	35.21	58.96
P K	0.07	0.06
Ca K	0.58	0.39
Ti K	39.66	22.18
Ru L	12.37	3.28
Ir M	5.69	0.79

Table 3. Ti/RuO₂·IrO₂ anode electrode coating composition (after use).

Spectrum	In stats.	C	O	Ca	Ti	Mn	Ru	Ir	Total
1	Yes	7.31			92.69				100.00
2	Yes	4.98	44.91	1.02	30.50	4.76	9.42	4.42	100.00
Max.		7.31	44.91	1.02	92.69	4.76	9.42	4.42	
Min.		4.98	44.91	1.02	30.50	4.76	9.42	4.42	

the cracked gaps, which could make for decreasing the Ohm drop and operating voltage in electrochemical reaction. Meanwhile, there was a bubble effect in the electrochemical reaction, and its influence could be expressed by two factors because: 1) the bubbles grew and adhered to the interface of the electrode and electrolyte, a bubble curtain was formed, which made the electrode active area decrease and allowed the electrode surface potential and current density to achieve microcosmic maldistribution and 2) the bubbles dispersed in solution, which made the electrolyte become a gas-liquid mixed system, resulting in real conductivity decreasing.

Meanwhile, Ohm drop and operating voltage increased and the energy consumption of the system was enhanced.

As seen in Fig. 8 and Tables 4 and 5, cathode electrode surface exhibited cracked characteristics to some extent before use, but it was not obvious, which may be owing to its manufacturing process. And it would not influence its electrochemical characteristics. Through the experiments, the change of cathode electrode surface morphology occurred, and the cracks gradually increased, which was owing to the redox reaction in the cathode electrode. In addition, the component of the cathode exhibited an obvious difference before and after use. Namely, Fe content decreased obviously, which could be considered as its consumption during redox reaction. Besides, quite a few fouling on the cathode electrode surface appeared after use, which would be formed with MgCO₃ and CaCO₃ in large measure. Therefore, the descaling should be paid close attention during practical application.

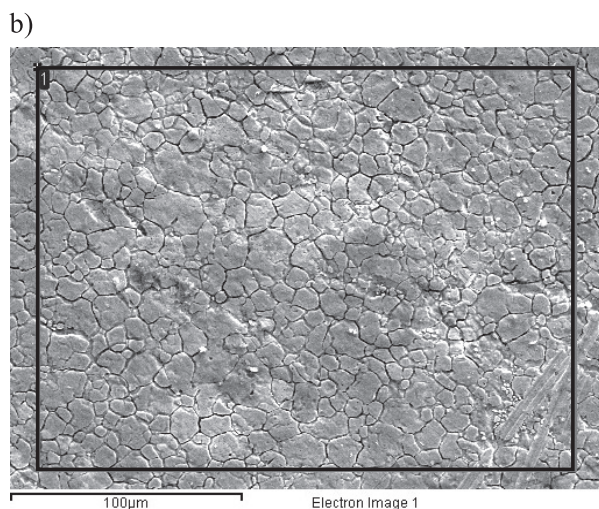
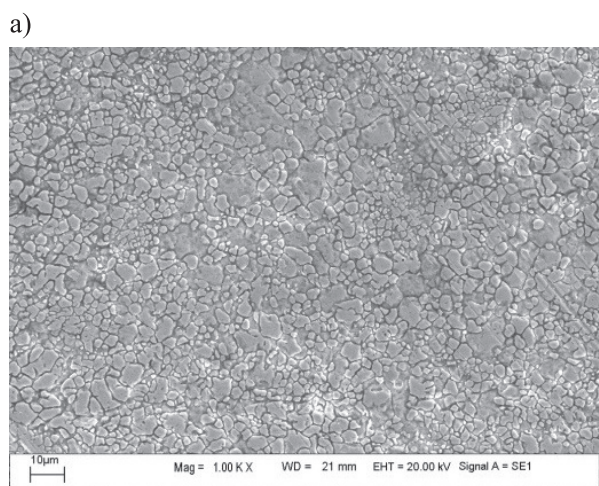


Fig. 7. Morphology characterization of cathode: a) before use, b) after use.

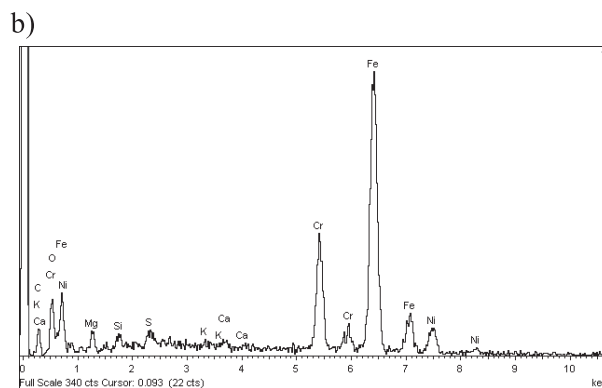
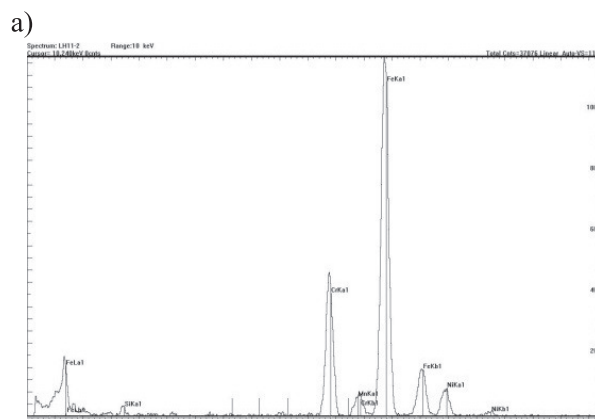


Fig. 8. Cathode EDS map: a) before use, b) after use.

Table 4. Cathode surface composition (before use).

Element	Line	Weight%	Error	K-Ratio	Cnts/s	Atomic%
Si	Ka	0.77	0.044	0.0039	6.00	1.51
Cr	Ka	17.13	0.20	0.19	138.09	18.11
S	Ka	0.74	0.049	0.0074	4.54	0.74
Fe	Ka	72.57	0.53	0.71	375.04	71.42
Ni	Ka	8.80	0.23	0.078	30.24	8.23

Table 5. Cathode surface composition (after use).

Element	Weight%	Atomic%
C K	13.31	35.76
O K	7.90	15.93
Mg K	2.02	2.68
Si K	0.79	0.91
S K	0.69	0.69
K K	0.21	0.17
Ca K	0.47	0.38
Cr K	13.41	8.33
Fe K	53.66	31.01
Ni K	7.54	4.15

Conclusions

In the experiments, we determined the operating optimum technological conditions using quite a few testing factors. The optimal parameters could be obtained for electrolysis time 60 min, current density 8 A, and electrode span 1 cm. This system could remove more than 73% of total organic carbon under optimum conditions. Through small pilot-scale continuous experiments and a scanning electron microscope, obvious change of element components on the electrodes surface were analyzed, especially Ti content on the anode and Fe content on the cathode. From the aforementioned analysis, the electrochemical method could work well in coking wastewater advanced treatment to remove organic compounds, which is difficult to be biodegraded.

Acknowledgements

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